

Products of electrolysis of dilute sulphuric acid

In the electrolysis of dilute sulphuric acid using platinum electrodes, the products obtained at the anode and cathode are: Anode Cathode A. Sulphur hydrogen D. Hydrogen oxygen No official explanation is available for this question at this time. Please check contributions posted by others below. If you can provide an explanation to help other student learn, please click here Report an Error Ask A Question Download App Tardigrade @. All rights reserved Electrolysis involves using electrolytes. The quantity of products of nead try again. Show answers Print Version (with answers) The passage of a direct current through an ionic compound in solution or molten form results in chemical reactions at the electrodes is related to the amount of electricity passed through the chemicals, and also to the number of moles of electrons required to form 1 mole of product. Task 1: HCI Electrolysis of dilute sulfuric acid. Task 1: HCI Electrolysis Model 1 Open Yenka is encoded. The simulation by the volume of gas collected in the inverted tube a sulton. At present, only water is contained in the apparatus. Two carbon rods act as electrodes, is open. The graph will show the volume of gas collected? (b) What volume of hydrogen is collected? (b) What volume of hydrogen is collected? (b) What volume of hydrogen is calculate the number of coulombs. The number of coulombs, a be converted to moles of electrons as the electrolysis? 1.35 × 10.3 moles of electrons is 24,000 cm3, calculate the number of moles of hydrogen produced by the jectrols with sequence by phydrogen as the form 7.08 × 10.4 moles of hydrogen is 0.4 moles of electrons erequired to run for the sequence as the product of the number of moles of electrons as the electroles as the electrol

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